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CHAPTER 10: Chemical
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ANSWER KEY

Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: **Chemical Quantities**

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SI unit representing
 6.02×10^{23}
representative
particles of a
substance. the
temperature and
pressure at which one
mole of gas occupies a
volume of 22.4 L. equal
volumes of gases at
the same temperature
and pressure contain
equal numbers of
particles.

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The presentation of information in this chapter assumes that you can already perform the tasks listed below. You can test your readiness to proceed by answering the Review Questions at the end of the chapter. This might also be a good time to read the Chapter Objectives, which precede the Review Questions.

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**Chapter 10 Chemical
Calculations and
equations**

Chapter 10 Chemical
Quantities 91 SECTION
10.1 THE MOLE: A
MEASUREMENT OF
MATTER (pages
287–296) This section
defines the mole and
explains how the mole
is used to measure
matter. It also teaches
you how to calculate

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the mass of a mole of
any substance.

Measuring Matter
(pages 287-289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287-296)

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... Chapter 11 -

Chemical Reactions

Chapter 12 -

Stoichiometry Chapter

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13 - States of Matter
Chapter 14 - Behavior
of Gases Chapter 15 -
Water and Aqueous
Systems Chapter 16 -
Solutions ...

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Chapter 10 (Chemical
Quantities) Test Study
Guide The mole is the
SI unit used to
measure the number of
representative
particles in a

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substance. A representative particle can be an atom, an ion, or a molecule, depending upon the way a substance commonly exists.

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