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## **Ap Chemistry Chapter 11 Practice**

A.P. Chemistry Practice Test: Ch. 11,  
Solutions Name\_\_\_\_\_ MULTIPLE CHOICE.

Choose the one alternative that best completes the statement or answers the question. 1) Formation of solutions where the process is endothermic can be spontaneous provided that \_\_\_\_\_.

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A)the solvent is a gas and the solute is a solid

## **A.P. Chemistry Practice Test: Ch. 11, Solutions MULTIPLE ...**

Chapter 11 AP Chemistry. gases. liquids. solids. converting a gas into a liquid or solid.... -highly compressible... -assumes shape of volume of container ... -m....

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- particles are not spread out ... -almost incompressible ... -assume....
- incompressible and have definite shape and volume ... -molecules....

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AP Chemistry Test (Chapter 11) Multiple Choice (50%) Please use the following

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choices to answer questions 1-7. A) London dispersion forces B) Ion-ion attractions C) Dipole-dipole attractions D) Dipole-ion attractions E) Hydrogen bonding H) Dipole-induced dipole attractions I) Ion-induced dipole attractions ...

## **AP Chemistry Test (Chapter 11)**

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AP Chemistry Chapter 11. Intermolecular Forces, Liquids, and Solids - 10 - Vapor Pressure and Boiling Point • Liquids boil when the external pressure at the liquid surface equals the vapor pressure. • The normal boiling point is the boiling point at 760 mm Hg (1 atm).

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## **Intermolecular Forces, Liquids ...**

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Chapter 11 – Properties of Solutions .  
11.1 Solution Composition . A. Molarity  
1. liters of. solution moles solute

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Molarity(M) =  $\frac{\text{mass of solute}}{\text{mass of solution}} \times 100 =$  Mass Percent  
percent. C. Mole Fraction =  $\frac{\text{moles of solute}}{\text{total moles of solution}}$   
D. Molality =  $\frac{\text{moles of solute}}{\text{kg of solvent}}$   
E. Normality =  $\frac{\text{equivalents of solute}}{\text{liter of solution}}$

## **Chapter 11 - Properties of Solutions**

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Ch 12: Ch 12 Practice Answers . Ch 13:  
Equilibrium Handout Key . Equilibrium  
Practice Answers . AP Question . Ch 14  
and 19: Ch 14 Ch19 Practice Keys .

**Baker, Mrs. (Science) / AP Chemistry**  
11%-15% of exam score . Unit 9:  
Applications of Thermodynamics You'll

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be introduced to the concept of “thermodynamic favorability” for reactions, meaning how likely they are to occur given energy changes and environmental factors. Topics may include: ... AP Chemistry Course and Exam Description

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energy.

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Chapter 6; Chapter 7; Chapters 8 & 9;  
Chapter 10; Chapter 11; Chapter 12;  
Chapter 13; Chapter 14; Chapter 15;  
Chapter 16; Chapter 17; Chapter 18;  
Chapter 22; Other; Labs; Chapter 11.

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Solutions. Formula ...

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## **Chapter 11 - Properties of Solutions | CourseNotes**

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created by me, but rather, were the collaborative effort of many AP Chemistry teachers. You may well find them useful. They only "work" correctly if you view them on a computer or tablet running PowerPoint. Chapter 2,3,7,8 PowerPoint Review (old Big Idea #1) handout Chapter 5,9,10 and 11 PowerPoint Review

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Time Limit None Allowed Attempts 2  
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At a given temperature the vapor pressure of pure liquid diethyl ether and isopropyl alcohol are 730 torr and 310 torr, respectively. A solution prepared by



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mixing isopropyl alcohol and diethyl ether obeys Raoult's law.

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Chemistry: Intermolecular Forces,  
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A Molecular Comparison of Liquids and  
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solids are due to intermolecular forces.

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