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SmileyKylie0923. Key Concepts: Terms in this set (57) Dalton. The atom is a tiny, indestructible particle with no internal structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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CHAPTER 5 What You'll Learn You will

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compare the wave and particle models of light. You will describe how the frequency of light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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Chapter 5: Electrons in Atoms - Neshaminy School District

Chapter 5 - Electrons in Atoms. Chapter
5 - Electrons in Atoms. Jennie L. Borders.
Section 5.1 - Models of the Atom. The
Rutherford's model of the atom did not
explain how an atom can emit light or
the chemical properties of an atom.
Plum Pudding Model Rutherford's Model.

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The Bohr Model.

Chapter 5 - Electrons in Atoms

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how: Aufbau principle - electrons enter the lowest energy first. This causes difficulties because of the overlap of orbitals of

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different energies - follow the diagram!
Pauli Exclusion Principle - at most 2
electrons per orbital ...

Chapter 5 Electrons in Atoms - Google Slides

Chapter 5 - electrons in atoms
(handouts) Chapter 6 - periodic table &
trends (handouts) Chapters 7/9 - ionic

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bonding & naming (handouts) Chapters
8/9 - covalent bonding & chemical
names & formulas (handouts) Chapters
8/15 - VSEPR/polar bonding/IMFs
(handouts) Chapter 10 - moles
(handouts) Chapter 11 - reactions
(handouts) Chapter 12 ...

Science / Chapter 5 - electrons in

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atoms (handouts)

Chapter 5: Electrons in Atoms Models of the Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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brinkley504. Ms. Cristina Chemistry I H.
Terms in this set (82) Rutherford's model-
also called nuclear model-did not explain

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how the electrons were arranged

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Chapter 5 - Electrons in Atoms. Chapter 5 - Electrons in Atoms. Section 5.1 - Models of the Atom. The Rutherford's model of the atom did not explain how an atom can emit light or the chemical

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properties of an atom. Plum Pudding Model Rutherford's Model. The Bohr Model.

Chapter 5 - Electrons in Atoms - CHEMISTRY with Crews

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136 Chapter 5 • Electrons in Atoms
Section 55.1.1 Figure 5.1 Different elements can have similar reactions with water. Objectives Compare the wave and particle natures of light. Define a quantum of energy, and explain how it is related to an energy change of matter. Contrast continuous electromagnetic

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Chapter 5: Electrons in Atoms

Chapter 5 Electrons in Atoms. Educators.
AY IB Chapter Questions. 02:25. Problem
1 Objects get their colors from reflecting
only certain wavelengths when hit with
white light. Light reflected from a green
leaf is found to have a wavelength of
 $4.90 \times 10^{-7} \text{ m}$.
What is the frequency of the light? ...

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How are electrons arranged in an atom?
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Chapter 5“Electrons in Atoms ...

Describe the energies and positions of
electrons according to the quantum
mechanical model. Section 5.1Models of

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the Atom. OBJECTIVES: Describe how the shapes of orbitals related to different sublevels differ. Ernest Rutherford's Model.

Chapter 5 Electrons in Atoms

Chapter 5 "Electrons in Atoms"

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How many electrons can each p orbital hold? Chapter 5: Electrons in Atoms

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