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Chapter 15: Acids and
Bases Acids and Bases
Arrhenius Definitions:
acids - compounds that
produce an increase in

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[H⁺] when dissolved in
water bases -

compounds that
produce an increase in
[OH⁻] when dissolved
in water Lewis

Definitions: acids -
electron pair acceptors
bases - electron pair
donors Brønsted-Lowry
Definitions: acids - H⁺
donors

Chapter 15: Acids and Bases Acids and Bases

Chapter 15 Review

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Acids Bases Answer

Key 3. solution. A
buffer solution is made
Introduction to pH,
pOH, and pKw
Autoionization of water
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is a.

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Questions 1) _____ is
found in carbonated
beverages due to the
reaction of carbon
dioxide with water.

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Bases, strong acid.

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Chemistry
strong base. weak acid.
weak base. an acid
that ionizes completely
in solvent. a base that
ionizes completely in a
solvent. an acid that
releases few hydrogen
ions in aqueous
solution. a base that
releases few hydroxide
ions in aqueous
solution.

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Acid-Base Indicator: A
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substance that changes color in a specific pH range. Indicators exhibit pH-dependent color changes because they are weak acids and have different colors in their acid (HIn) and conjugate base (In⁻) forms. $\text{HIn(aq)} + \text{H}_2\text{O} \rightleftharpoons \text{In}^- + \text{H}_3\text{O}^+$

Chapter 15 Equilibria: Acids and Bases

Chapter 15 Acids and
Bases, STUDY. PLAY.

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Chapter 15 main idea.
acids and bases
change the color of
compound indicators.
... -Lewis acid-base
reaction is the
formation of one or
more covalent bonds
between an electron-
pair donor and an
electron pair acceptor,
although the 3 acid-
base reactions differ,
many compounds can
be categorized as acids
...

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Chapter 15 Acids and Bases

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ACIDS AND BASES 453

SECTION 15-1 O

OBJECTIVES List five general properties of aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and

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base accord-ing to

Arrhenius's theory of
ionization. Explain the
differences

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Bases 1. The
hydronium ion and the

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CHAPTER 14 REVIEW
Acids and Bases

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SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1.

Name the following compounds as acids:
sulfuric acid a. H_2SO_4
sulfurous acid b. H_2SO_3
hydrosulfuric acid c. H_2S
perchloric acid d. HClO_4
hydrocyanic acid e. hydrogen cyanide 2. H_2CO_3

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Chapter 14 & 15 . HW:
•Read Ch 14: •Fill in as
much of the acid base
table as you can, as
you read . Acid base
conductivity and
reactivity Conductivity,
Reac'vity,, ... 3.Work
on ch 14-15 review
(due Wednesday, test
Friday!) • Solutions in
which [H

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Chapter 14 & 15
Chapter 15 Review
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CHAPTER 15 REVIEW
Acids and Bases -
Calendar CHAPTER 15
REVIEW Acids and
Bases SECTION 15-1
SHORT ANSWER

Answer the following
questions in the space
provided

1 Name the
following compounds
as acids: a H_2SO_4 b
 H_2SO_3 c H_2S d
 HClO_4 e hydrogen
cyanide

2 Which (if
any) of the acids
mentioned in item 1

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**[PDF] Chapter 15
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2 Answers**

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that is a compound of
hydrogen, oxygen, and
a third el.... a
substance that
increases the
concentration of
HYDRONIUM ions.... a
substance that

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increases the concentration of HYDROXIDE ions.... an acid that is a compound of hydrogen, oxygen, and a third el....

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Zumdahl, Susan A. ,
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Chapter 14 - Acids and Bases - Review Questions - Page 699: 2

Aside from a mention
of a textbook page,
this chemistry
assignment is a
suitable review of acids

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Chemistry
and bases. It begins by addressing conjugate pairs and acid base reactions.

Neutralization and amphoteric properties are also dealt with. The worksheet concludes with two real-life word problems.

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